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Periodic Table With Chemistry Formulas SparkCharts

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Periodic Table with Chemistry Formulas

CHEMISTRY FORMULAS

FORMS AND SATURATIONS PARTICLES

- A. Atom.** The smallest particle of an element that retains the properties of that element. Atoms are collectively known as matter.
- B. Molecule.** The smallest particle of matter consisting of two or more atoms, measured in atomic mass units (amu). A molecule is a group of atoms held together by chemical bonds.
- C. Ions.** Atoms (ions) of an element that have lost or gained one or more electrons, measured in atomic mass units (amu).

- D. Isotopes.** Atoms (ions) of an element that have the same number of protons but different numbers of neutrons.
- E. Atoms.** $A = Z + N$
- F. Molecules.** $M = Z_1 + Z_2 + \dots + Z_n$
- G. Ions.** $I = Z + \Delta Z$
- H. Atoms.** $A = Z + N$
- I. Molecules.** $M = Z_1 + Z_2 + \dots + Z_n$
- J. Isotopes.** Atoms (ions) of an element that have the same number of protons but different numbers of neutrons.
- K. Atoms.** $A = Z + N$
- L. Molecules.** $M = Z_1 + Z_2 + \dots + Z_n$
- M. Ions.** $I = Z + \Delta Z$

Name	Symbol	Mass	Charge
Proton	p^+	1 amu	+1
Neutron	n^0	1 amu	0
Electron	e^-	1.67×10^{-24} amu	-1

- N. Molecular mass.** The mass of a molecule, measured in atomic mass units (amu).
- O. The atomic mass of an element equals the mass of one atom of that element.**
- P. The mass of atoms (ions) of an element equals the mass of one atom of that element.**

- Q. MOLECULAR MASS**
- R. Average molecular mass.** $M = Z_1 + Z_2 + \dots + Z_n$
- S. The atomic mass of an element equals the mass of one atom of that element.**
- T. The mass of atoms (ions) of an element equals the mass of one atom of that element.**

MOLECULAR FORMULA AND EMPirical FORMULA

- A. Molecular formula.** The chemical formula of a compound that gives the actual number of atoms of each element in a molecule of the compound.
- B. Empirical formula.** The chemical formula of a compound that gives the relative number of atoms of each element in a molecule of the compound.
- C. Isotopes.** Atoms (ions) of an element that have different numbers of neutrons.
- D. The empirical formula for element X is X_2 .**

BALANCED CHEMICAL EQUATIONS

- A. Check for elements.** Check the chart on the left.
- B. Balance words.**
- C. Balance numbers.**
- D. Balance atoms.**
- E. Balance hydrogen.**
- F. You can always take coefficients from the left and subtract them at all other places.**
- G. You can always take coefficients from the right and subtract them at all other places.**
- H. Divide the equation by their greatest common divisor.**

PERCENTAGE AND MOLARITY

- A. Percentage composition.** tells how much mass of each element is contained in 100 g of a compound.
- B. Molar mass.** $M = Z_1 + Z_2 + \dots + Z_n$
- C. Molar mass.** $M = \frac{1000}{\text{percent}}$
- D. Parts per million.** $\text{ppm} = \frac{\text{mass of element}}{\text{mass of a solution}} \times 10^6$
- E. Molar concentration.** $C = \frac{\text{moles of solute}}{\text{liters of solution}}$

INTERACTIONS OF MATTER INTERACTIONS

- A. Attraction.** The tendency of particles to attract one another.
- B. Dispersion.** $\text{London dispersion} = \frac{1}{2} \times \frac{Z_1 Z_2}{r^6}$
- C. Dipole.** $\text{Dipole-dipole} = \frac{Z_1 Z_2}{r^3}$

- D. Ion-dipole.** $\text{Ion-dipole} = \frac{Z_1 Z_2}{r^3}$
- E. Ion-ion.** $\text{Ion-ion} = \frac{Z_1 Z_2}{r^3}$

- F. Hydrogen bonding.** $\text{Hydrogen bonding} = \frac{Z_1 Z_2}{r^3}$
- G. Ion-pair.** $\text{Ion-pair} = \frac{Z_1 Z_2}{r^3}$

IDEAL GAS LAWS

- A. Pressure.** $P = \frac{F}{A}$
- B. Boyle's law.** $P_1 V_1 = P_2 V_2$
- C. Charles's law.** $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
- D. Gay-Lussac's law.** $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

- E. Avogadro's law.** $\frac{V_1}{N_1} = \frac{V_2}{N_2}$
- F. Ideal gas law.** $PV = NRT$

- G. Henry's law.** $P = K_c N$
- H. Raoult's law.** $P = P_0 X$

- I. Dalton's law of partial pressure.** $P = P_1 + P_2 + P_3 + \dots$
- J. Partial pressure.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$

- K. Henry's law constant.** $K_c = \frac{P}{N}$
- L. Raoult's law constant.** $K = \frac{P}{N}$

- M. Partial pressure of a gas.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$
- N. Partial pressure of a gas.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$

- O. Henry's law.** $P = K_c N$
- P. Partial pressure of a gas.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$

- Q. Raoult's law.** $P = P_0 X$
- R. Dalton's law of partial pressure.** $P = P_1 + P_2 + P_3 + \dots$

- S. Henry's law.** $P = K_c N$
- T. Raoult's law.** $P = P_0 X$

- U. Henry's law constant.** $K_c = \frac{P}{N}$
- V. Raoult's law constant.** $K = \frac{P}{N}$

- W. Henry's law.** $P = K_c N$
- X. Raoult's law.** $P = P_0 X$

- Y. Dalton's law of partial pressure.** $P = P_1 + P_2 + P_3 + \dots$
- Z. Partial pressure.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$

- A. Henry's law.** $P = K_c N$
- B. Raoult's law.** $P = P_0 X$

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- G. Partial pressure of a gas.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$
- H. Partial pressure of a gas.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$

- I. Henry's law.** $P = K_c N$
- J. Raoult's law.** $P = P_0 X$

- K. Dalton's law of partial pressure.** $P = P_1 + P_2 + P_3 + \dots$
- L. Partial pressure.** $P = \frac{N}{N_{\text{tot}}} P_{\text{tot}}$

- M. Boiling point.** The temperature at which a liquid begins to boil.
- N. Boiling point elevation.** $\Delta T_b = K_b \times m$
- O. Boiling point depression.** $\Delta T_d = K_d \times m$
- P. Boiling point.** $T_b = T_0 + \Delta T_b$
- Q. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- R. Boiling point elevation.** $T_b = T_0 + \Delta T_b$
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- Y. Boiling point elevation.** $T_b = T_0 + \Delta T_b$
- Z. Boiling point.** $T_b = T_0 + \Delta T_b$

- A. Boiling point.** $T_b = T_0 + \Delta T_b$
- B. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- C. Boiling point.** $T_b = T_0 + \Delta T_b$
- D. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- E. Boiling point.** $T_b = T_0 + \Delta T_b$
- F. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- G. Boiling point.** $T_b = T_0 + \Delta T_b$
- H. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- I. Boiling point.** $T_b = T_0 + \Delta T_b$
- J. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- K. Boiling point.** $T_b = T_0 + \Delta T_b$
- L. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- M. Boiling point.** $T_b = T_0 + \Delta T_b$
- N. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- O. Boiling point.** $T_b = T_0 + \Delta T_b$
- P. Boiling point depression.** $T_b = T_0 - \Delta T_b$
- Q. Boiling point elevation.** $T_b = T_0 + \Delta T_b$
- R. Boiling point.** $T_b = T_0 + \Delta T_b$
- <b



Synopsis

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